# Application Notes

> **RE: Table of Tabulation**
> **Appendix A**
> (Tab 33)

## Application Notes

## Water vs. Fossil-Fuel Energy Content

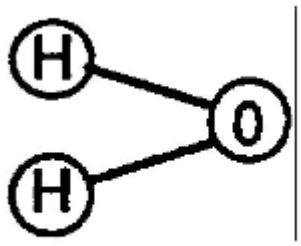
Water is composed of (2) Hydrogen Atoms and (1) Oxygen Atom to form a molecule of Water.

**Atomic Mass Unit:**

Electron (E) = 1 Proton (P) - 1Mu Hydrogen
Atom: IE = IP - IMu Oxygen Atom: 8E = 8P - 8Mu Atomic Mass Ratio (Mur) of Water
(2H X 1Mu) plus (1 Oxy. X 8 Mu) = 10 Mu's
**See Appendix (B) Note (2)**

Whereby,
2H (Mu) divided by (10 Mu's) = 20%

Thus,
One gallon of Water contains 1.6691 lbs. of Hydrogen




**Molecular Structure of Water**
(Volumetric Displacement of Atom spheres)

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**Energy-Yield Potential of Water**

One water gallon equals 8.345 lbs

8.345 lbs x .20 = 1.669 pounds of Hydrogen / $H_2O$ gal.

1.669 pounds of hydrogen-fuel of water - .183591 lbs (11% per volume of impurities ... typically 20 ppm - 40 ppm contaminates with Ambient Air being present) =

1.4854 lbs of hydrogen atoms available for gas combustion per gallon of Water approximately.

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**Water as Fuel**
The by-product of burning gases derived from Water is environmentally safe since there is no . *UOOJIS* present in the Water molecule ... resulting in the re-formation of Water "mist" after gas combustion... being able to re-energize the newly formed Water Droplets for energy "reuse" once exposed to Sunlight. **(See Energy recycling graph 530 of Figure 5-6, once again)**

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**Appx. A 01**

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Revision #2
Created 2024-10-15 17:20:23 UTC by Chris Bake
Updated 2024-10-15 17:25:56 UTC by Chris Bake