

# Application Notes

**RE: Table of Tabulation**  
**Appendix A**  
(Tab 33)

## Application Notes

### Water vs. Fossil-Fuel Energy Content

Water is composed of (2) Hydrogen Atoms and (1) Oxygen Atom to form a molecule of Water.

#### **Atomic Mass Unit:**

Electron (E) = 1 Proton (P) - 1Mu Hydrogen

Atom:  $1E = 1P - 1\text{Mu}$  Oxygen Atom:  $8E = 8P - 8\text{Mu}$  Atomic Mass Ratio (Mur) of Water

$(2H \times 1\text{Mu}) \text{ plus } (1 \text{ Oxy.} \times 8 \text{ Mu}) = 10 \text{ Mu's}$

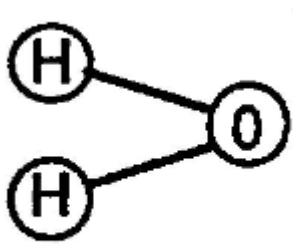
**See Appendix (B) Note (2)**

Whereby,

$2H (\text{Mu}) \text{ divided by } (10 \text{ Mu's}) = 20\%$

Thus,

One gallon of Water contains 1.6691 lbs. of Hydrogen



#### **Molecular Structure of Water**

(Volumetric Displacement of Atom spheres)

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#### **Energy-Yield Potential of Water**

One water gallon equals 8.345 lbs

$8.345 \text{ lbs} \times .20 = 1.669 \text{ pounds of Hydrogen / H}_2\text{O gal.}$

1.669 pounds of hydrogen-fuel of water - .183591 lbs (11% per volume of impurities ... typically 20 ppm - 40 ppm contaminates with Ambient Air being present) =

1.4854 lbs of hydrogen atoms available for gas combustion per gallon of Water approximately.

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### **Water as Fuel**

The by-product of burning gases derived from Water is environmentally safe since there is no . *UOOJIS* present in the Water molecule ... resulting in the re-formation of Water "mist" after gas combustion... being able to re-energize the newly formed Water Droplets for energy "reuse" once exposed to Sunlight. **(See Energy recycling graph 530 of Figure 5-6, once again)**

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**Appx. A 01**

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