

Application Notes

RE: Table of Tabulation
Appendix A
(Tab 33)

Application Notes

Water vs. Fossil-Fuel Energy Content

Water is composed of (2) Hydrogen Atoms and (1) Oxygen Atom to form a molecule of Water.

Atomic Mass Unit:

Electron (E) = 1 Proton (P) - 1Mu Hydrogen

Atom: $1E = 1P - 1\text{Mu}$ Oxygen Atom: $8E = 8P - 8\text{Mu}$ Atomic Mass Ratio (Mur) of Water

$(2H \times 1\text{Mu}) \text{ plus } (1 \text{Oxy.} \times 8 \text{Mu}) = 10 \text{ Mu's}$

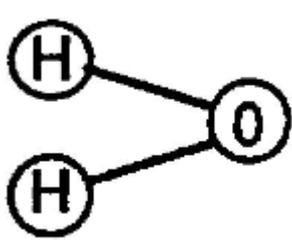
See Appendix (B) Note (2)

Whereby,

$2H (\text{Mu}) \text{ divided by } (10 \text{ Mu's}) = 20\%$

Thus,

One gallon of Water contains 1.6691 lbs. of Hydrogen



Molecular Structure of Water

(Volumetric Displacement of Atom spheres)

Energy-Yield Potential of Water

One water gallon equals 8.345 lbs

$8.345 \text{ lbs} \times .20 = 1.669 \text{ pounds of Hydrogen / H}_2\text{O gal.}$

1.669 pounds of hydrogen-fuel of water - .183591 lbs (11% per volume of impurities ... typically 20 ppm - 40 ppm contaminates with Ambient Air being present) =

1.4854 lbs of hydrogen atoms available for gas combustion per gallon of Water approximately.

Water as Fuel

The by-product of burning gases derived from Water is environmentally safe since there is no . *UOOJIS* present in the Water molecule ... resulting in the re-formation of Water "mist" after gas combustion... being able to re-energize the newly formed Water Droplets for energy "reuse" once exposed to Sunlight. **(See Energy recycling graph 530 of Figure 5-6, once again)**

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Appx. A 01

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