Atomic Interaction to Voltage Stimulation

Atomic structure of an atom exhibits two types of electrical charged mass-entities.

Orbital electrons having **negative electrical charges** (-) and a **nucleus** composed of **protons** having **positive electrical charges** (+).

In stable electrical state, the number of negative electrically charged electrons **equals** the same number of positive electrically charged protons ... forming an atom having "no" **net electrical charge**.

> Whenever one or more electrons are "**dislodged**" from the atom, the atom takes on a **net positive electrical charge** and is called a **positive ion**.

> If an electron **combines** with a stable or normal atom, the atom has a **net negative charge** and is called a **negative ion**.

Voltage potential within an electrical circuit (see **Voltage Intensifier Circuit** as to Figure 1-1) can cause one or more electrons to be dislodged from the atom due to opposite polarity attraction between unlike charged entities, as shown in Figure (1-8).

Voltage Intensifier Circuit as to Figure 1-1

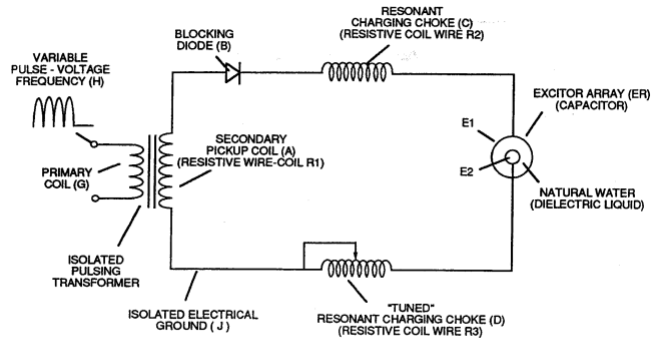

FIGURE 1-1: VOLTAGE INTENSIFIER CIRCUIT (AA)

Figure (1-8)

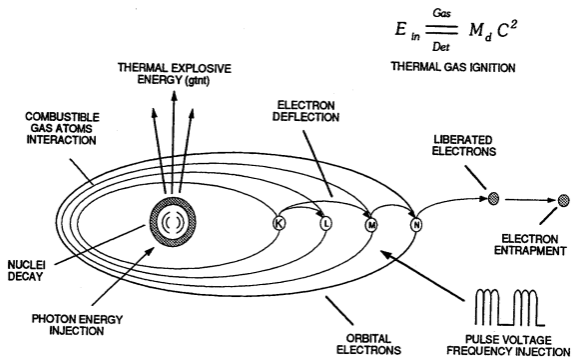

FIGURE 1-8: HYDROGEN FRACTURING PROCESS

> (see Figure 1-6 again as to Figure 1-9) as to **Newton's** and **Coulomb's Laws** of **electrical force** (RR).

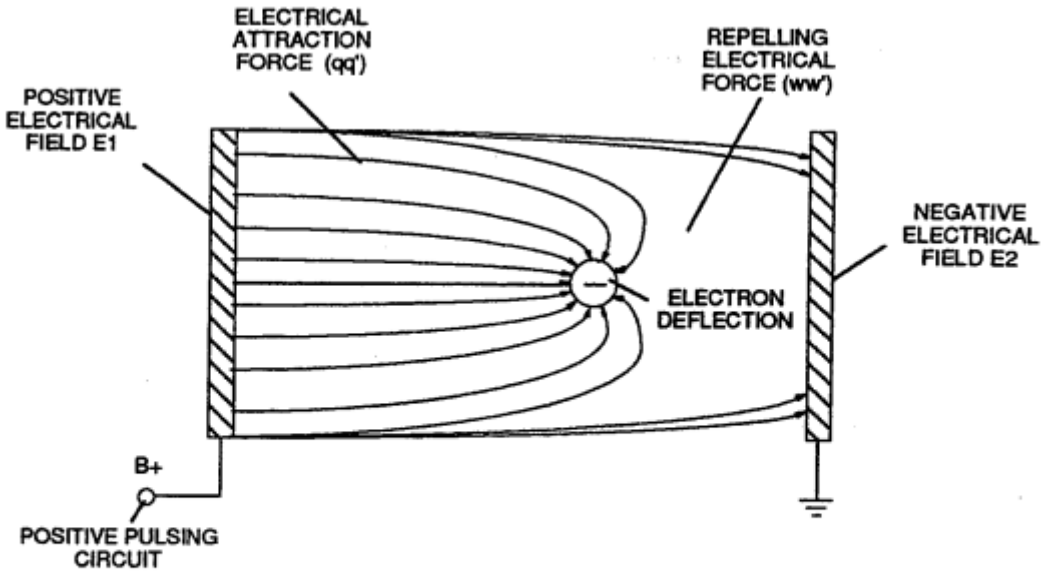

FIGURE 1-6: VOLTAGE POTENTIAL PERFORMING WORK

> The resultant electrical attraction force (qq') combines or joins unlike atoms together by way of **covalent bonding** to form molecules of gases, solids, or liquids.

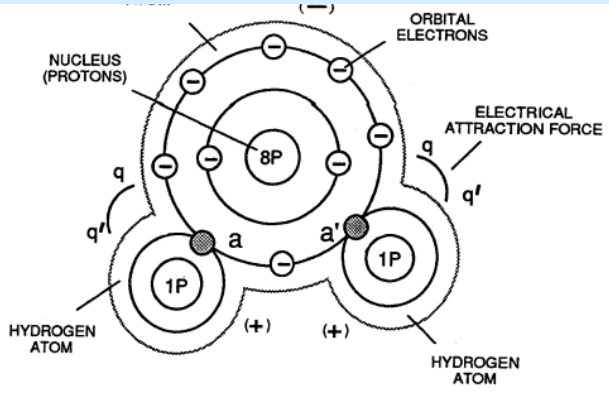

FIGURE 1-7: ELECTRICAL CHARGES OF THE WATER MOLECULE

When the unlike oxygen atom combines with two hydrogen atoms to from the water molecule by accepting the hydrogen electrons (aa' of Figure 1-7), the oxygen atoms become "net" **negative electrically charged** (-) since the restructured oxygen atom now occupies 10 negative electrically charged electrons as to only 8 positive electrically charged protons.

The hydrogen atom with only its **positive charged proton remaining** and **unused**, now, takes on a "net" **positive electrical charge** *equal to* the electrical intensity of the negative charges of the two **electrons** (aa') being shared by the oxygen atom.

... satisfying the law of physics that *for every action there is an equal and opposite reaction*.

The sum total of the **two positive charged hydrogen atoms** (++) equaling the **negative charged oxygen atom** (--) forms a "no" net electrical charged molecule of water.

> Only the unlike atoms of the water molecule exhibits opposite electrical charges.

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