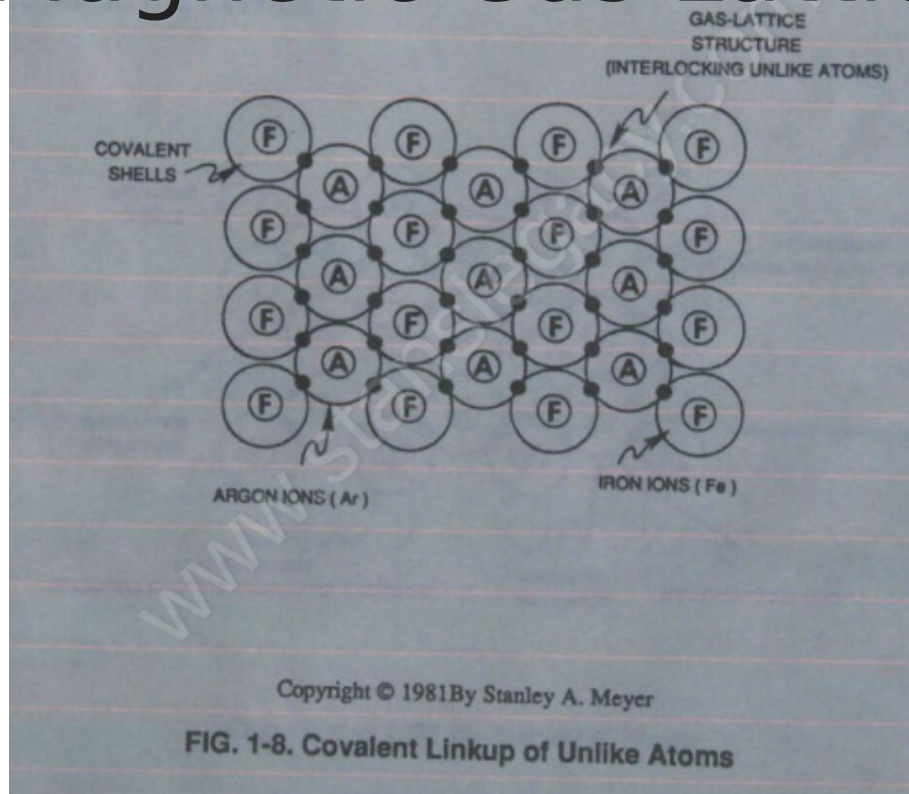


# Magnetic Gas Lattice



The forming **Argon ion** ( $\text{Ar}^+$ ) is

now exposed to **Iron ions** ( $\text{Fe}^+$ ) (magnetic properties) experiencing and undergoing the same **Electron Extraction Process**.

Together, the **two ions** ( $\text{Ar}^+/\text{Fe}^+$ ) form a covalent link up or **covalent bond** when the covalent electron of the **Argon ion** ( $\text{Ar}^+$ ) pair up and be shared with the valence electron of the **Iron ion** ( $\text{Fe}^+$ ).

Covalent bonding of **Iron ions** ( $\text{Fe}^+$ ) to the **Argon ion** ( $\text{Ar}^+$ ) continues until a geometrical **Gas-Lattice Structure** is formed, as illustrated in Figure 1-8.

Stable-state of the **Gas-Lattice** occurs when the covalent shell of each unlike atom structure becomes full or filled up... the **Argon atom** (Ar) sees an covalent shell of 8 electrons while, at the same time, the **Iron atom** (Fe) sees an **covalent shell** (M shell) of 14 electrons.

Covalent bonding between like atoms does not occur due to the "stronger" **Electrical Attraction-Force** ( $qq'$ ) between the unlike atoms.

During Gas-Lattice formation, **Iron ions** ( $\text{Fe}^+$ ) can be replaced by other atoms exhibiting magnetic properties such as **Nickel ions** ( $\text{Ni}^+$ ) or **Cobalt ions** ( $\text{Co}^+$ ).

**Gas-Lattice formation** of unlike atoms by way of the **Electron Extraction Process** is, hereinafter, called "**The Gas Bonding Process**".

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